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P3-15 Assuming the reactions to be elementary: $2 \text{ Anthracene} \rightarrow \text{Dimer}$ $\frac{dC_D}{dt} = k_f C_A^2 - k_r C_D$ where, $K_c = \frac{k_f}{k_r}$ Similarly for the second reaction: $3 \text{ Norbornadiene} \rightarrow \text{Quadricyclane}$ $\frac{dC_Q}{dt} = k_f C_N^3 - k_r C_Q$ where, $K_c = \frac{k_f}{k_r}$ P3-16 The mistakes are as follows: 1. For any reaction, the rate law cannot be written on the basis of the stoichiometric equation.

Solution Manual Essentials of Chemical Reaction ...

Stoichiometry. We shall set up Stoichiometric Tables using A as our basis of calculation in the following reaction. We will use the stoichiometric tables to express the concentration as a function of conversion. We will combine $C_i = f(X)$ with the appropriate rate law to obtain $-r_A = f(X)$.

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